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### 6.4 a, b, c Subatomic Particles

The table below contains information about several elements. In each case, enough information has been provided for you to fill in the blanks. Assume all atoms are neutral (have no charge).

Name	Symbol	Atomic Number	Atomic Mass Number (rounded)	# of Protons	# of Electrons	# of Neutrons
1. Calcium	Ca	20	40	20	20	20
2. Magnesium	Mg	12	24	12	12	12
3. Hydrogen	H	1	1	1	1	0
4. Gold	Au	79	197	79	79	118
5. Iron	Fe	26	56	26	26	30
6. Mercury	Hg	80	201	80	80	121
7. Chlorine	Cl	17	35	17	17	18
8. Bismuth	Bi	83	209	83	83	126
9. Strontium	Sr	38	88	38	38	50

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**Isotopes**

The number of protons that an atom has is the atomic number. The number of neutrons that an atom has is the mass number. The number of electrons that an atom has is equal to the number of protons.

Example: Carbon has 6 protons and 6 electrons. It has 6 neutrons. Its mass number is 12.

Carbon-12 has 6 protons and 6 neutrons. Its mass number is 12.

Carbon-13 has 6 protons and 7 neutrons. Its mass number is 13.

Carbon-14 has 6 protons and 8 neutrons. Its mass number is 14.

Carbon-12 and carbon-13 are isotopes of carbon. They have the same number of protons but different numbers of neutrons.

Isotopes of an element have the same chemical properties but different physical properties.

Isotopes are used in many ways, such as in medicine, industry, and research.

Example: Carbon-14 is used to date ancient objects.

Example: Iodine-131 is used to treat thyroid cancer.

Example: Uranium-235 is used as fuel in nuclear reactors.

Example: Cobalt-60 is used in cancer treatment.

Example: Strontium-90 is used in cancer treatment.

Example: Technetium-99m is used in medical imaging.

Example: Radium-226 is used in cancer treatment.

Example: Polonium-210 is used in cancer treatment.

Example: Astatine-211 is used in cancer treatment.

Example: Francium-223 is used in cancer treatment.

Example: Actinium-227 is used in cancer treatment.

Example: Thorium-232 is used in cancer treatment.

Example: Uranium-238 is used in cancer treatment.

Example: Plutonium-239 is used in cancer treatment.

Example: Americium-241 is used in cancer treatment.

Example: Curium-244 is used in cancer treatment.

Example: Berkelium-247 is used in cancer treatment.

Example: Californium-251 is used in cancer treatment.

Example: Einsteinium-252 is used in cancer treatment.

Example: Fermium-257 is used in cancer treatment.

Example: Mendelevium-258 is used in cancer treatment.

Example: Nobelium-259 is used in cancer treatment.

Example: Lawrencium-260 is used in cancer treatment.

Example: Rutherfordium-261 is used in cancer treatment.

Example: Dubnium-262 is used in cancer treatment.

Example: Seaborgium-263 is used in cancer treatment.

Example: Bohrium-264 is used in cancer treatment.

Example: Hassium-265 is used in cancer treatment.

Example: Meitnerium-266 is used in cancer treatment.

Example: Darmstadtium-267 is used in cancer treatment.

Example: Roentgenium-268 is used in cancer treatment.

Example: Copernicium-269 is used in cancer treatment.

Example: Nihonium-270 is used in cancer treatment.

Example: Flerovium-271 is used in cancer treatment.

Example: Tennessine-272 is used in cancer treatment.

Example: Oganesson-273 is used in cancer treatment.

### Determine the Number of Protons, Neutrons, & Electrons

Use the Periodic Table of Elements to complete the following chart

Chemical Name	Chemical Symbol	Atomic Number	Mass Number	Number of Protons	Number of Electrons	Number of Neutrons
Hydrogen	H	1	1	1	1	0
Boron	B	5	11	5	5	6
Carbon	C	6	12	6	6	6
Beryllium	Be	4	9	4	4	5
Helium	He	2	4	2	2	2
Lithium	Li	3	7	3	3	4
Nitrogen	N	7	14	7	7	7
Neon	Ne	10	20	10	10	10
Fluorine	F	9	19	9	9	10
Oxygen	O	8	16	8	8	8
Sodium	Na	11	23	11	11	12
Aluminum	Al	13	27	13	13	14
Magnesium	Mg	12	24	12	12	12
Silicon	Si	14	28	14	14	14
Phosphorus	P	15	31	15	15	16
Argon	Ar	18	40	18	18	22
Chlorine	Cl	17	36	17	17	19
Sulfur	S	16	32	16	16	16
Potassium	K	19	39	19	19	20
Calcium	Ca	20	40	20	20	20
Iron	Fe	26	56	26	26	30
Copper	Cu	29	64	29	29	35

**Note** the mass number of an element is the atomic mass rounded to a whole number

objectives of the unit: I can distinguish between protons, neutrons and electrons I/4 I can determine the number of protons, neutrons and electrons in neutral atoms, ions and isotopes I/4 I can read the periodic table and explain its layout, groups and system of classification system include various sections, such as multiple choice, "who are," draw atoms and more, neutrons, electrons, isotopes, cations, anions-30+ examples of atom, isotope and ion analysis in terms of mass number, atomic number and subatomic particles this work is authorized by the creative-non-noder-noder-derivative attribution license 4.0. to display a copy of this license, visit 3one editable worksheet can be used as preview, revision, evaluation, etc. matter, atom, nucleus, proton, neutron, electron, spectral line, energy level, electron, element, atomic number, concepts of understanding1) illustrates an atom, including protons, electrons and neutrons. 2) compares the mass, position and charge of the three basic particles of an atom. 3) lists the three observations of Dalton. 4) illustrated Thomson's, Rutherford's, anpage 4periodic table handout for students to practice finding information of the periodic table: symbols, atomic number, atomic mass, protons, neutrons, electrons, metals, metalloids, nonmetals, solids, liquids, gas, page 5covers scientific method; definitions of matter, mass, volume, density; density calculation, mass, volume; measurements (righello, graduated cylinder, triple radius balance, thermometer; form/volume/particles for gas, liquid and solids; protons, neutrons and electrons; history of the periodic table; identify the element by symbol, atomicAtomic mass, group and period. Mainly the multiple choice with a little filling of the empty game, sublimation, molecule, solubility, chemical change, reactivity, atom, proton, electron, neutron, nucleus, atomic number, compound. Page 7 This easy game to play is a great way to practice essential biological concepts. This PDF contains words concerning the base of life chemistry for the Biology classes I. Words include: ATP, nucleic acids, proteins, lipids, carbohydrates, catalyst, activation energy, enzymes, pH, acid, base, proton, neutron, ions, electrons of value and more on short video page with this free product, your students will love coloring, cutting and assembling these 3D atomic models of the first 10 elements to create a beautiful and lasting display for your class. Each atomic model is provided with its own information tag including element symbol, owners and uses, atomic number, atomic mass and much more. From this information tags, students determine the number of protons, neutrons, electrons and the electrons configuration for each Atopage 9seter B Unit VI guideline with activator questions; Objectives

and a general temporal sequence on lessons. The terms can include: electromagnetic waves, radio waves, infrared rays, ultraviolet rays, X-rays, gamma rays, wavelength, frequency, spectrum, black body radiation, doppler effect, shift of the redshift / blueshift, transversal and lines of emission of radial speed / enthusiastic states, photons or how much nucleus, radiation area, convection area, photosphere, chromosphere, transition area, live work sheets> English finish !! It is to allow access to the microphone look at the upper part of the web browser. If you see a message asking for authorization to access the .weiverp .weiverp siht ni nwohs ton si 2 egaP weiverP eerF a gnidaeR er'uoY esolC .wolla esaelp

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